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the inert gas structure of 8-electrons in their outermost shell by the sharing of electrons. Since Carbon atoms can achieve the inert gas electron arrangement only by the sharing of electrons, therefore, Carbon always forms covalent bonds.

MIC  
ZIT  
MAD  
MAD

## Carbon and its Compounds

Carbon always forms covalent bonds. For this we know that the atomic no. of Carbon is 6 which means that a neutral atom of Carbon contains 6 electrons. So the electronic configuration of Carbon is  $K_2 L_4$ . Thus it is clear that Carbon has 4 electrons in its outer shell. So it should either lose 4 electrons or gain 4 electrons to achieve the inert gas electronic configuration and become stable. Now Carbon atom is very small due to which its outermost electrons are strongly held by the nucleus. So, it is not possible to remove 4 electrons from a Carbon atom. It is also not possible to add as many as 4 electrons to a Carbon atom due to energy considerations - and acquire

Co-ordination - Chemistry

Simple double and Complex Salts:

Complex Salts → A Complex Salt is a Salt

that contains one or more ~~Geo~~ Complex ions:

ions with metal centers and different

molecules attached

In chemistry, a salt is any compound

composed of oppositely charged ions. These ions

are negative (called anions) and positive

(called cation), resulting in a charge of

zero. Complex Salts are distinguished from

simple salts and double salt which have

different makeups.