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Lewis concept of Acid and bases →

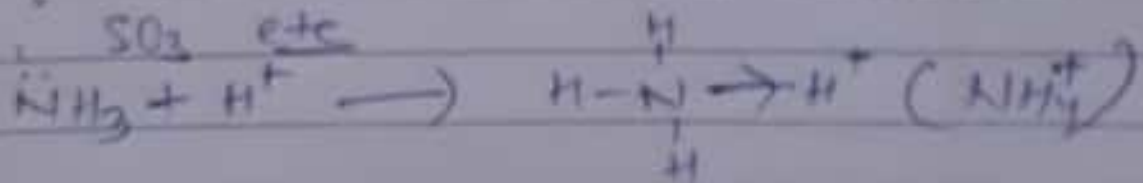
Arrhenius and Bronsted-Lowry concepts of Acid and bases are limited to hydrogen and oxygen containing substances are acid and base. But G.N. Lewis in 1923 generalised the term of acid and bases, which is based on electrons. therefore Lewis concept is also called electronic concept of acids and bases.

According to this concept —

Acid is a substance (an atom, molecule or ion) which can accept a pair of electrons and base is a substance (an atom, molecule or ion) which can donate a pair of electrons.

ie Electron pair acceptor = Acid
" " donor = Base.

example of Lewis - H^+ , NH_4^+ , NO_2^+ , $C_2H_5^+$, BF_3 , $AlCl_3$, $FeCl_3$, SO_2 , SO_3 etc



The neutralization ~~is~~ between Acid and Base means the formation of co-ordinate bond

Base: → Acid

