



Although the two bonds are the equivalent to the original symmetric molecule, the bond-dissociation energy of an oxygen-hydrogen bond varies slightly depending on whether or not there is another hydrogen atom bonded

Somewhat depending on the specific molecule, so tabulated bond energies are generally averages from a number of selected typical chemical species containing that type of bond. Bond energy (BE) is the average of all bond-dissociation energies of a single type of bond in a given molecule. The bond-dissociation energies of several different bonds of the same type can vary even within a single molecule. For example, a water molecule is composed of two O-H bonds bonded as H-O-H. The bond energy for H_2O is the average of energy required