

is liberated at the anode

2. Whytlaw Gray's Method: → For this method pure dry and anhydrous potassium hydrogen fluoride is electrolyzed in the molten state at 523K. when hydrogen is liberated at the cathode while F_2 is liberated at the anode.

Physical properties of Fluorine:

↳ Colour: ~~Fluorine~~ fluorine is a pale yellow gas
Odour: It has an irritating pungent odour
Nature: It is ~~is~~ highly ~~is~~ poisonous and has a corrosive action on skin

Density: It is heavier than air.

Boiling point: 86 K

Preparation of Fluorine

The major method for preparing fluorine is electrolytic oxidation. The most common electrolysis procedure is to use a molten mixture of potassium hydrogen fluoride, KHF_2 and anhydrous hydrogen fluoride. Electrolysis causes HF to decompose forming fluorine gas at the anode and hydrogen at the cathode.

Here we can describe some of important method for preparation of fluorine.

Dennis Method :- The principle of this method is pure dry and anhydrous potassium hydrogen fluoride (KHF_2) is electrolyzed in the molten state at 523 K when hydrogen is liberated at the cathode while F_2 is