

of ozone may go up 20-25%.

Principle of preparation of ozone:

In the reaction, 142.7 kJ of energy

is absorbed for the formation of one mole of ozone.

The reaction is initiated by a sparkless or

silent electric discharge, to produce less

heat, as ozone is prone to decomposing back

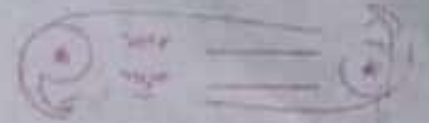
into oxygen with a rise in temperature.

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OZONE Part 1 (H)

Ozone (or tri-oxygen) is a triatomic molecule with the chemical formula O_3 . It is a pale blue gas with a distinctively pungent smell. It is an allotrope of oxygen that is much less stable than the diatomic allotrope O_2 , breaking down in the lower atmosphere.

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A fairly low temperature (around 0°C) should be maintained.

There should be no sparking.

⊗ Brodie's ozoniser

In principle, this ozoniser is like the Siemen's ozoniser but dilute sulphuric