

# Hybridization

Part 1 (H) & (Subj)

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In chemistry hybridization is the concept of mixing atomic orbitals into new hybrid orbitals suitable for the pairing of electrons to form chemical bonds in valence bond theory. Hybrid orbitals are very useful in the explanation of molecular geometry and atomic bonding properties and are symmetrically disposed in space. Although sometimes taught together with the valence shell electron-pair repulsion theory, valence bond and hybridisation are in fact not related to the VSEPR model.

Hybridisation theory is an integral part of organic chemistry, one of the most compelling examples being, Baldwin's rules for drawing reaction mechanisms. Sometimes

with two atoms sharing 4 two  
Hybridisation theory explains bonding  
in alkenes and methane. The amount  
of p character or s character, which is  
decided mainly by orbital hybridisation  
can be used to reliably predict molecu