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Bond energy

In chemistry, bond energy (BE), also called the mean body enthalpy or average bond. Enthalpy is the measure of bond strength in a chemical bond. IUPAC defines bond energy as the average value of the gas-phase bond-dissociation energy (usually at a temperature of 298.15K) for all bonds of the same type within the same chemical species. The larger the average bond energy, per electron pair bond, of a molec

enthalpy change of the following fission: $R-x \rightarrow R + x$. The BDE, denoted by $D^\circ(R-x)$, is usually derived by the thermochemical equation:

$$D^\circ(R-x) = \Delta H_f(R) + \Delta H_f(x) -$$

The enthalpy of formation $\Delta H f^\circ$ of a large number of atoms, free radicals, ions, clusters and compounds is available from the websites of NIST, NASA, CODATA and IUPAC. Most

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Formic acid, due to the electron withdrawing effect of chlorine, promoting ionization in benzoic acid, the carbon atoms which are present in the ring are sp^2 hybridized. As a result, benzoic acid is a stronger acid than cyclohexane carboxylic acid, due to aromatic