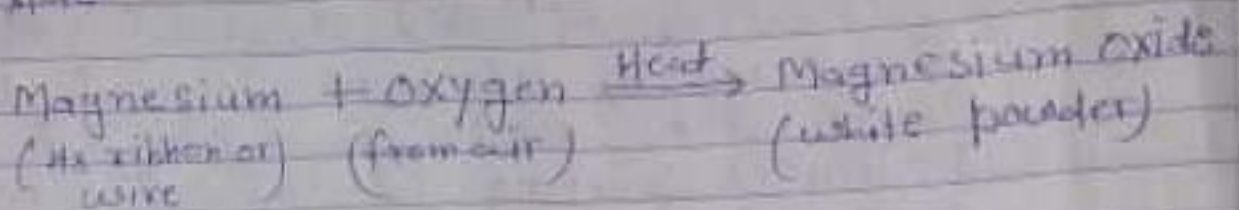


flame to form a white powder called magnesium oxide. Actually on heating, magnesium combines with oxygen present in air to form magnesium oxide.



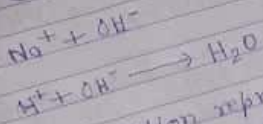
have The burning of Magnesium in air to form magnesium oxide is an example of chemical reaction. In this chemical reaction there are two reactants one is magnesium and other is oxygen but only one product that is magnesium oxide.

Chemical reactions can be carried not only in Science laboratory but a large no. of chemical reaction keep on occurring in our daily life such as Souring of milk.

Formation of curd from milk, cooking of food, Digestion of food in our body, Burning of fuels, Ripening of fruit etc.

Kamlesh Kumar

NaOH



Since the reaction representing neutralisation process involving the combination of H^+ and OH^- ions. The approximately constant molar heat of neutralisation would be expected. Thus the constant heat of

neutralisation of a strong acid by a strong base is readily understandable in terms of this concept. This concept

has offered a means of correlating catalytic behaviour with the concentration of the H^+ ions.

Kaushikendra Kumar

17/4

Base Part 1 (H)

Physical chemistry

Chapter 2

Modern Concept of acid & Base \rightarrow In general word we can say that ~~the~~ Acids are those chemical substances which have a sour taste and acid change the colour of blue litmus to red.

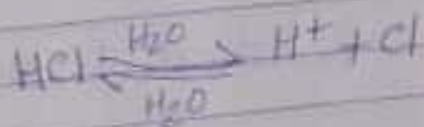
Bases are those chemical substances which have a bitter taste all the bases changes the colour of red litmus to blue.

But the modern concept is difference that is

According to this concept an acid is any hydrogen containing compound which gives H^+ ions in aqueous solution and a base

which gives OH^- ions in aqueous solution. For example \rightarrow The HCl is an acid and NaOH is a base and the neutralisation

process can be represented by a reaction involving the combination of H^+ and OH^- ions to form H_2O



17/4

CHEMICAL REACTIONS

1950 Part 1 (Sub)

In general way we can say that chemical reaction involve chemical changes. During chemical reactions a rearrangement of atoms takes place between the reacting substances to form new substances having entirely different properties. It also involving breaking of old chemical bonds and then making of new chemical bonds between the rearranged atoms of new substances. During a chemical reaction atoms of one element do not changes into those of another element.

So we can define the chemical reactions is such type \rightarrow "Chemical reactions are the processes in which new substances with new properties are formed".

The substances which take part in a chemical reaction are called reactants and the new substances produced as a result of chemical reaction are called products. The products thus formed have properties which are entirely different from those of the reaction.

Example of a chemical reaction

In Science laboratory -

when a magnesium ribbon or new magnesium wire is heated, it burns in air with a dazzling white