

classmate
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is dissolved in a solvent other than water, it is called a non aqueous solution. eg → Iodine in Carbon tetrachloride, Sulfur in Carbon disulfide, Phosphorus in ethyl alcohol

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Basic

BSc. (Part 1)

Type of Solutions

There are different types of solutions which can be classified on such type that is

1 Supersaturated Solution → Comprises of a large amount of solute at a temperature wherein it will be reduced as a result that the extra solute will crystallize quickly.

2 Unsaturated Solution → is a solution in which a solvent is capable of dissolving any more solute at a given temperature.

3 Saturated Solution → can be defined as a solution in which a solvent is not capable of dissolving any more solute at a given temperature.

4 Aqueous Solution → when a solute is dissolved in water the solution is called an aqueous solution.

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Classification of Solid

on the basis of the nature of order present in the arrangement of constituent particles solids can be classified as crystalline or amorphous.

Crystalline Solids -

A crystalline solid is a substance whose constituent particles possess a regular orderly long range arrangement.

Example \Rightarrow Sodium chloride, sucrose (can sugar), diamond, quartz, copper, potassium nitrate etc.

A crystalline solid usually consist of a large no of small crystals.

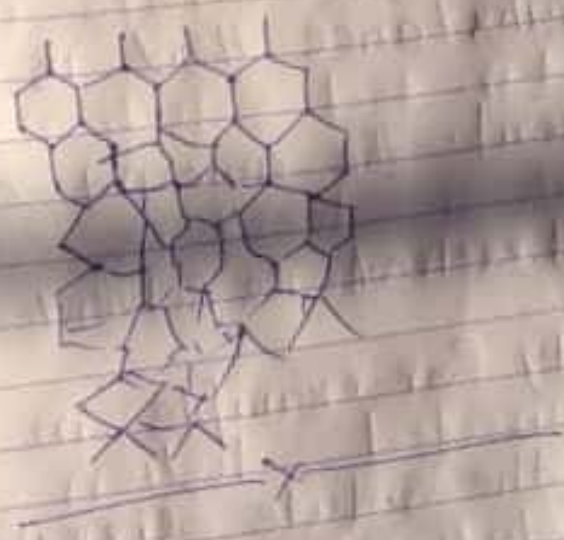
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Solution

Base (part 1)

In generally we can say that a solution is a homogeneous mixture of two or more substance. A solution may exist in any phase. A solution consists of a solute and a solvent. The solute is the substance that is dissolved in the solvent. The amount of solute that can be dissolved in solvent is called its solubility. For example, in a saline solution salt is the solute dissolved in water as the solvent. For solution with components in the same phase, the substances present in lower concentration are solute, while the substance present in highest abundance

each of them having a definite characteristic geometrical shape. The arrangement of constituent particles is ordered and repeats itself periodically over the entire crystal.



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y The molecules of Solids are fixed at one one point.

z Intra particles distances are short.

z Most of the Solids are hard incompressible and rigid.